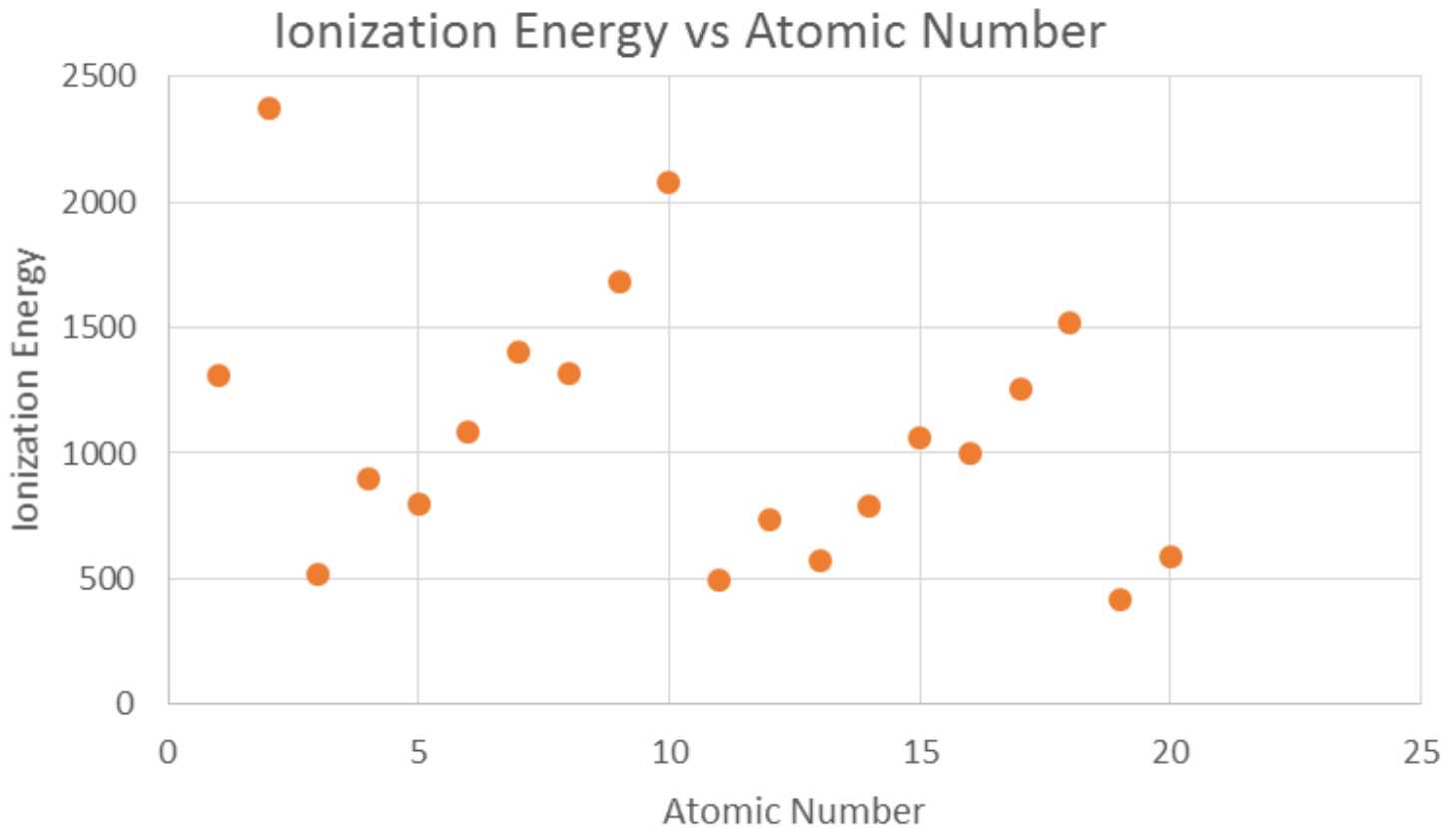


The Graph above is a diagram showing the relationship between the atomic radius and atomic number of the first twenty elements. The atomic radius of an element is a measure of the size of its atoms defined by the distance from the nucleus to the boundary of the surrounding electron cloud.

As you can see from the graph above the atomic radius decreases as you move across a period as shown by the declining line of points. On the contrary as you move down a period the atomic radius increases as you can see by the height of the peaks or top points of the decline that the size of the atoms are increasing.



The graph above is a representation of the trend ionization energy of the first twenty elements in the periodic table. Ionization energy is the ionization energy of an atom or molecule describes the amount of energy required to remove an electron from the atom or molecule.

Periodic law states that as you move down a group the ionization energy increases as you move down a group as shown by the positive slope of the line of points. Also according to periodic law the ionization energy decreases as you move across a period as shown by the decreasing peaks in the line3s of the graph.